



General Chemistry (Second Semester)

This course covers material usually taught in the second semester of a two-semester sequence. By default, the topics listed below are all available. However, instructors can customize the course to align with their teaching goals using *any* topics from the complete ALEKS curriculum, which covers a full course sequence.

Curriculum Show All (287 topics + 546 additional topics)

- Math and Physics (28 topics)
 - ◆ Mathematics (6 topics)
 - ◇ Integer multiplication and division
 - ◇ Simplifying a fraction
 - ◇ Equivalent fractions
 - ◇ Writing expressions using exponents
 - ◇ Evaluating expressions with exponents of zero
 - ◇ Square root of a perfect square
 - ◆ Algebra Expressions (4 topics)
 - ◇ Using distribution and combining like terms to simplify: Univariate
 - ◇ Quotient rule with negative exponents: Problem type 1
 - ◇ Power and quotient rules with positive exponents
 - ◇ Square root of a perfect square monomial
 - ◆ Linear Equations (6 topics)
 - ◇ Additive property of equality with integers
 - ◇ Multiplicative property of equality with signed fractions
 - ◇ Solving a linear equation with several occurrences of the variable: Fractional forms with monomial numerators
 - ◇ Solving a rational equation that simplifies to linear: Denominator $x+a$
 - ◇ Solving for a variable in terms of other variables using addition or subtraction with division
 - ◇ Solving for a variable in terms of other variables in a linear equation with fractions
 - ◆ Quadratic and Radical Equations (3 topics)
 - ◇ Solving a quadratic equation using the square root property: Decimal answers, basic
 - ◇ Solving a quadratic equation using the square root property: Decimal answers, advanced
 - ◇ Applying the quadratic formula: Decimal answers
 - ◆ Graphing Equations (3 topics)
 - ◇ Classifying slopes given graphs of lines
 - ◇ Graphing a line through a given point with a given slope
 - ◇ Finding slope given the graph of a line on a grid
 - ◆ Logarithms and Exponentials (2 topics)
 - ◇ Evaluating a logarithmic expression
 - ◇ Solving an equation of the form $\log_b a = c$
 - ◆ Electrostatics (4 topics)
 - ◇ Understanding that opposite charges attract and like charges repel
 - ◇ Understanding net electrical charge
 - ◇ Understanding how electrostatic forces cancel
 - ◇ Sketching polarization induced by a nearby charge
- Measurement and Matter (12 topics)
 - ◆ Scientific Notation (1 topics)

- ◇ Converting between decimal numbers and numbers written in scientific notation
- ◆ SI Units (2 topics)
 - ◇ Knowing the value of an SI prefix as a power of 10
 - ◇ Interconverting temperatures in Celsius and Kelvins
- ◆ Measurement Uncertainty (3 topics)
 - ◇ Counting significant digits
 - ◇ Counting significant digits when measurements are added or subtracted
 - ◇ Counting significant digits when measurements are multiplied or divided
- ◆ Quantitative Problem Solving (4 topics)
 - ◇ Predicting the units of the solution to a basic quantitative problem
 - ◇ Deducing the unit missing from the solution to a basic quantitative problem
 - ◇ Setting up the math for a one-step problem with unit conversion
 - ◇ Setting up the math for a two-step quantitative problem
- ◆ Chemical Elements (2 topics)
 - ◇ Recognizing element families
 - ◇ Organization of the Periodic Table
- Atoms, Ions and Molecules (5 topics)
 - ◆ Atomic Structure (2 topics)
 - ◇ Predicting the ions formed by common main-group elements
 - ◇ Counting valence electrons in a neutral atom
 - ◆ Chemical Compounds (1 topics)
 - ◇ Understanding the prefixes used in naming binary compounds
 - ◆ Ionic Compounds (2 topics)
 - ◇ Deducing the ions in a binary ionic compound from its empirical formula
 - ◇ Naming ionic compounds with common oxoanions
- Chemical Reactions (16 topics)
 - ◆ Moles and Molar Mass (1 topics)
 - ◇ Finding molar mass from chemical formulae
 - ◆ Chemical Equations (3 topics)
 - ◇ Balancing chemical equations with interfering coefficients
 - ◇ Writing a chemical equation from a molecular model
 - ◇ Writing the net equation for a sequence of reactions
 - ◆ Reaction Stoichiometry (3 topics)
 - ◇ Using a chemical equation to find moles of product from moles of reactant
 - ◇ Solving for a reactant using a chemical equation
 - ◇ Identifying the limiting reactant in a drawing of a mixture
 - ◆ Solution Stoichiometry (2 topics)
 - ◇ Calculating molarity using solute mass
 - ◇ Using molarity to find solute mass and solution volume
 - ◆ Precipitation (2 topics)
 - ◇ Identifying the correct sketch of a compound in aqueous solution
 - ◇ Writing net ionic equations
 - ◆ Acid-Base Reactions (2 topics)
 - ◇ Predicting the products of a neutralization reaction
 - ◇ Determining the volume of base needed to titrate a given mass of acid
 - ◆ Oxidation-Reduction Reactions (2 topics)
 - ◇ Assigning oxidation numbers
 - ◇ Recognizing reduction and oxidation
 - ◆ Predicting Products (1 topics)
 - ◇ Predicting products from a general statement about reactivity
- Thermochemistry (3 topics)
 - ◆ Energy (2 topics)
 - ◇ Understanding the definitions of heat and work

- ◇ Understanding the definition of enthalpy
 - ◆ Reaction Enthalpy (1 topics)
 - ◇ Using the general properties of reaction enthalpy
- Electronic Structure and Chemical Bonding (24 topics)
 - ◆ Quantum Mechanics (1 topics)
 - ◇ Recognizing s and p orbitals
 - ◆ Electron Configuration (3 topics)
 - ◇ Deducing n and l from a subshell label
 - ◇ Drawing a box diagram of the electron configuration of an atom
 - ◇ Writing the electron configuration of an atom using the Periodic Table
 - ◆ Electronic Properties of The Elements (6 topics)
 - ◇ Identifying s, p, d and f block elements
 - ◇ Identifying elements with a similar valence electron configuration
 - ◇ Understanding periodic trends in effective nuclear charge
 - ◇ Deducing the block of an element from an electron configuration
 - ◇ Understanding periodic trends in atomic size
 - ◇ Understanding periodic trends in atomic ionizability
 - ◆ Lewis Structures (5 topics)
 - ◇ Counting valence electrons in a molecule or polyatomic ion
 - ◇ Calculating formal charge
 - ◇ Writing Lewis structures for a molecule with one central atom and no octet–rule exceptions
 - ◇ Writing the Lewis structures for a molecule with resonance
 - ◇ Drawing Lewis structures for simple organic compounds
 - ◆ Chemical Bonds (3 topics)
 - ◇ Predicting the relative electronegativities of atoms
 - ◇ Predicting relative bond polarity
 - ◇ Predicting the relative length and energy of chemical bonds
 - ◆ Molecules (4 topics)
 - ◇ Predicting bond angles in a small organic molecule
 - ◇ Predicting whether molecules are polar or nonpolar
 - ◇ Naming common chemical groups
 - ◇ Identifying common chemical groups in a Lewis structure
 - ◆ VB and MO Theory (2 topics)
 - ◇ Counting sigma and pi bonds in a small molecule
 - ◇ Identifying carbon hybridization in simple organic molecules
- States of Matter (39 topics)
 - ◆ Gas Laws (1 topics)
 - ◇ Using the ideal equation of state
 - ◆ Intermolecular Forces (6 topics)
 - ◇ Identifying a molecule from its electrostatic potential map
 - ◇ Predicting the strength of intermolecular forces from an electrostatic potential map
 - ◇ Identifying hydrogen–bonding interactions between molecules
 - ◇ Identifying the intermolecular forces between atoms, ions and molecules
 - ◇ Predicting the relative strength of the dispersion force between molecules
 - ◇ Predicting the relative boiling points of pure substances
 - ◆ Liquids (4 topics)
 - ◇ Identifying important physical properties of liquids
 - ◇ Understanding consequences of important physical properties of liquids
 - ◇ Relating vapor pressure to vaporization
 - ◇ Understanding the connection between vapor pressure, boiling point, and enthalpy of vaporization
 - ◆ Solids (7 topics)
 - ◇ Predicting the type of solid formed by a compound
 - ◇ Predicting the relative stability of ionic crystals from a sketch

- ◇ Predicting the relative lattice energy of binary ionic compounds
- ◇ Drawing the unit cell of a 2D lattice
- ◇ Counting the atoms in a unit cell
- ◇ Recognizing and naming close-packed crystal lattices
- ◇ Recognizing and naming lattices with cubic unit cells
- ◆ Phase Change (7 topics)
 - ◇ Using heat of fusion or vaporization to find the heat needed to melt or boil a substance
 - ◇ Using a phase diagram to predict phase at a given temperature and pressure
 - ◇ Labeling a typical simple phase diagram
 - ◇ Using a phase diagram to find a phase transition temperature or pressure
 - ◇ Sketching a described thermodynamic change on a phase diagram
 - ◇ Identifying phase transitions on a heating curve
 - ◇ Interpreting a heating curve
- ◆ Solution Composition (5 topics)
 - ◇ Calculating mass percent composition
 - ◇ Calculating molality
 - ◇ Calculating mole fraction
 - ◇ Calculating mass concentration
 - ◇ Solving applied dilution problems
- ◆ Solubility (2 topics)
 - ◇ Applying like dissolves like
 - ◇ Understanding how solubility varies with temperature and pressure
- ◆ Ideal Solutions (7 topics)
 - ◇ Predicting relative boiling point elevations and freezing point depressions
 - ◇ Using the K_f and K_b equations
 - ◇ Using the K_f and K_b equations with electrolytes
 - ◇ Using osmotic pressure to find molar mass
 - ◇ Using a solution freezing point to calculate a molar mass
 - ◇ Using Raoult's Law to calculate the vapor pressure of a component
 - ◇ Calculating ideal solution composition after a distillation
- Kinetics and Equilibrium (49 topics)
 - ◆ Rates of Reaction (3 topics)
 - ◇ Predicting how reaction rate varies with pressure, concentration and temperature
 - ◇ Calculating the reaction rate of one reactant from that of another
 - ◇ Calculating average and instantaneous reaction rate from a graph of concentration versus time
 - ◆ Rate Laws (11 topics)
 - ◇ Using a rate law
 - ◇ Using reactant reaction order to predict changes in initial rate
 - ◇ Deducing a rate law from initial reaction rate data
 - ◇ Calculating the change in concentration after a whole number of half-lives of a first-order reaction
 - ◇ Using a zero order integrated rate law to find concentration change
 - ◇ Using an integrated rate law for a first-order reaction
 - ◇ Using a second-order integrated rate law to find concentration change
 - ◇ Using first- and second-order integrated rate laws
 - ◇ Deducing a rate law from the change in concentration over time
 - ◇ Finding half life and rate constant from a graph of concentration versus time
 - ◇ Solving applied problems with first-order kinetics
 - ◆ Activation Energy (4 topics)
 - ◇ Interpreting a reaction energy diagram
 - ◇ Relating activation energy to reaction rate
 - ◇ Drawing the reaction energy diagram of a catalyzed reaction
 - ◇ Understanding the qualitative predictions of the Arrhenius equation
 - ◆ Reaction Mechanisms (8 topics)

- ◇ Identifying the molecularity of an elementary reaction
- ◇ Identifying intermediates in a reaction mechanism
- ◇ Writing a plausible missing step for a simple reaction mechanism
- ◇ Writing the rate law of an elementary reaction
- ◇ Writing the rate law implied by a simple mechanism with an initial slow step
- ◇ Expressing the concentration of an intermediate in terms of the concentration of reactants
- ◇ Writing the rate law implied by a simple mechanism
- ◇ Deducing information about reaction mechanisms from a reaction energy diagram
- ◆ Dynamic Equilibrium (4 topics)
 - ◇ Understanding that no reaction goes to 100% completion
 - ◇ Predicting relative forward and reverse rates of reaction in a dynamic equilibrium
 - ◇ Using Le Chatelier's Principle to predict the result of changing concentration
 - ◇ Using Le Chatelier's Principle to predict the result of changing temperature
- ◆ Equilibrium Constants and Expressions (7 topics)
 - ◇ Writing a concentration equilibrium constant expression
 - ◇ Writing a pressure equilibrium constant expression
 - ◇ Writing the concentration equilibrium expression for a heterogeneous equilibrium
 - ◇ Writing the pressure equilibrium expression for a heterogeneous equilibrium
 - ◇ Calculating an equilibrium constant from an equilibrium composition
 - ◇ Using an equilibrium constant to predict the direction of spontaneous reaction
 - ◇ Using the general properties of equilibrium constants
- ◆ Equilibrium Composition (7 topics)
 - ◇ Recognizing equilibrium from a sketch
 - ◇ Predicting equilibrium composition from a sketch
 - ◇ Setting up a reaction table
 - ◇ Calculating equilibrium composition from an equilibrium constant
 - ◇ Calculating an equilibrium constant from a partial equilibrium composition
 - ◇ Calculating an equilibrium composition after a prior equilibrium determines K
 - ◇ Solving problems that mix equilibrium ideas with gas laws
- ◆ Ionic Equilibria (5 topics)
 - ◇ Writing a solubility product (K_{sp}) expression
 - ◇ Using K_{sp} to calculate the solubility of a compound
 - ◇ Using the solubility of a compound to calculate K_{sp}
 - ◇ Calculating the solubility of an ionic compound when a common ion is present
 - ◇ Understanding the effect of pH on the solubility of ionic compounds
- Acids and Bases (43 topics)
 - ◆ Concepts of Acidity (11 topics)
 - ◇ Identifying acids and bases by their reaction with water
 - ◇ Understanding the difference between strong and weak acids
 - ◇ Identifying Bronsted–Lowry acids and bases
 - ◇ Identifying strong or weak acids and bases from a sketch
 - ◇ Finding the conjugate of an acid or base
 - ◇ Predicting acid or base strength from the conjugate
 - ◇ Predicting the products of the reaction of a strong acid with water
 - ◇ Predicting the reactants of a neutralization reaction
 - ◇ Predicting the qualitative acid–base properties of salts
 - ◇ Predicting the qualitative acid–base properties of metal cations
 - ◇ Identifying Lewis acids and bases in reactions
 - ◆ Acid–Base Nomenclature (4 topics)
 - ◇ Naming inorganic acids
 - ◇ Deducing the formulae of inorganic acids from their names
 - ◇ Naming acid salts
 - ◇ Recognizing common acids and bases

- ◆ Acidity and Chemical Structure (2 topics)
 - ◇ Predicting the relative acidity of binary acids
 - ◇ Understanding the effect of induction on acidity
- ◆ pH (8 topics)
 - ◇ Interconverting pH and hydronium ion concentration
 - ◇ Interconverting pH and pOH at 25°C
 - ◇ Interconverting hydronium and hydroxide concentration at 25°C
 - ◇ Making qualitative estimates of pH change
 - ◇ Calculating the pH of a strong acid solution
 - ◇ Calculating the pH of a strong base solution
 - ◇ Diluting a strong acid solution to a given pH
 - ◇ Preparing a strong base solution with a given pH
- ◆ Weak Acids and Bases (11 topics)
 - ◇ Writing an acid dissociation constant expression
 - ◇ Determining the strength of acids from a sketch
 - ◇ Calculating the K_a of a weak acid from pH
 - ◇ Calculating the pH of a weak acid solution
 - ◇ Writing a base protonation constant expression
 - ◇ Calculating the pH of a weak base solution
 - ◇ Deriving K_b from K_a
 - ◇ Interconverting K_a and pKa
 - ◇ Calculating the pH of a salt solution
 - ◇ Calculating percent dissociation of a weak acid
 - ◇ Writing the dissociation reactions of a polyprotic acid
- ◆ Acid–Base Titrations (3 topics)
 - ◇ Calculating the pH of a weak acid titrated with a strong base
 - ◇ Calculating the pH of a weak base titrated with a strong acid
 - ◇ Calculating the pH at equivalence of a titration
- ◆ Buffers (4 topics)
 - ◇ Identifying the major species in weak acid or weak base equilibria
 - ◇ Setting up a reaction table for a pH calculation with a common ion
 - ◇ Calculating the pH of a buffer
 - ◇ Calculating the composition of a buffer of a given pH
- Entropy and Free Energy (15 topics)
 - ◆ Entropy (8 topics)
 - ◇ Calculating entropy change from reversible heat flow
 - ◇ Calculating absolute entropy using the Boltzmann hypothesis
 - ◇ Calculating entropy change using the Boltzmann hypothesis
 - ◇ Predicting qualitatively how entropy changes with temperature and volume
 - ◇ Predicting qualitatively how entropy changes with mixing and separation
 - ◇ Qualitatively predicting reaction entropy
 - ◇ Using the Second Law to predict spontaneous change
 - ◇ Calculating reaction entropy using the standard molar entropies of reactants
 - ◆ Gibbs Free Energy (7 topics)
 - ◇ Using the general properties of Gibbs free energy
 - ◇ Calculating dG from dH and dS
 - ◇ Using the conditions of spontaneity to deduce the signs of H and S
 - ◇ Interconverting standard Gibbs free energy and K
 - ◇ Using thermodynamic data to calculate K
 - ◇ Recognizing consistency between statements about standard Gibbs free energy
 - ◇ Using the maximum work theorem with chemical work
- Electrochemistry (15 topics)
 - ◆ Half Reactions (5 topics)

- ◇ Writing a simple half–reaction from its description
- ◇ Writing and balancing complex half–reactions in acidic solution
- ◇ Writing and balancing complex half–reactions in basic solution
- ◇ Balancing a complex redox equation in acidic or basic solution
- ◇ Writing the half–reactions of a complex redox reaction in acidic or basic solution
- ◆ Galvanic Cells (8 topics)
 - ◇ Designing a galvanic cell from two half–reactions
 - ◇ Analyzing a galvanic cell
 - ◇ Picking a reduction or oxidation that will make a galvanic cell work
 - ◇ Ranking the strength of oxidizing and reducing agents using standard reduction potentials
 - ◇ Calculating standard reaction free energy from standard reduction potentials
 - ◇ Recognizing consistency among equilibrium constant, free energy, and cell potential
 - ◇ Using the Nernst equation to calculate nonstandard cell voltage
 - ◇ Understanding concentration cells
- ◆ Electrolysis (2 topics)
 - ◇ Using the relationship between charge, current and time
 - ◇ Using the Faraday constant
- Chemistry of the Elements (13 topics)
 - ◆ Periodic Trends (6 topics)
 - ◇ Understanding main–group periodic trends in metallicity
 - ◇ Predicting the most positive and negative oxidation states of main–group elements
 - ◇ Predicting the common oxidation states of main–group elements
 - ◇ Identifying a main–block group from an element oxide
 - ◇ Identifying a main–block group from an element halide
 - ◇ Predicting the type of bonding in a main–group element
 - ◆ Properties of Groups (4 topics)
 - ◇ Predicting the products of the reaction of a Group 1A or 2A metal with water
 - ◇ Predicting the products of the reaction of a Group 1A or 2A metal with oxygen
 - ◇ Predicting the products of the reaction of elements at either end of the Periodic Table
 - ◇ Ordering the melting points of elements at either end of the Periodic Table
 - ◆ Electronic Structure of Transition Metals (2 topics)
 - ◇ Writing the electron configuration of a first transition series atom
 - ◇ Identifying transition metal cations with a given number of d electrons
 - ◆ Properties of Transition Metals (1 topics)
 - ◇ Understanding words that describe where transition metals lie in the Periodic Table
- Nuclear Chemistry (14 topics)
 - ◆ Nuclear Reactions (5 topics)
 - ◇ Interpreting the symbol for a nuclide
 - ◇ Writing the symbols in a nuclear chemical equation
 - ◇ Balancing a nuclear chemical equation
 - ◇ Writing the equation for a typical radioactive decay
 - ◇ Calculating the energy change in a nuclear reaction from the mass change
 - ◆ Radioactive Decay (5 topics)
 - ◇ Knowing the properties of the common types of nuclear radiation
 - ◇ Understanding the common modes of radioactive decay
 - ◇ Understanding radioactive half life
 - ◇ Interconverting amount of radioactive decay and half life
 - ◇ Calculating radioactive activity from half life
 - ◆ Nuclear Technology (4 topics)
 - ◇ Using isotope ratios to radiodate
 - ◇ Using activity to radiodate
 - ◇ Knowing units of radiation dosage and exposure
 - ◇ Identifying important types of nuclear medicine procedure

- Organic Chemistry (11 topics)
 - ◆ Physical Properties of Organic Compounds (2 topics)
 - ◇ Identifying organic compounds
 - ◇ Identifying rigid parts of an acyclic organic molecule
 - ◆ Drawing Organic Molecules (2 topics)
 - ◇ Interpreting condensed chemical structures
 - ◇ Interpreting condensed chemical structures with benzene rings
 - ◆ Principles of Organic Nomenclature (5 topics)
 - ◇ Naming normal alkanes
 - ◇ Identifying the main chain of branched alkanes
 - ◇ Naming the parent hydrocarbon of branched alkanes
 - ◇ Identifying organic functional groups
 - ◇ Using family suffixes to name organic compounds
 - ◆ Organic Isomerism (2 topics)
 - ◇ Identifying constitutional isomers
 - ◇ Drawing the condensed structure of a constitutional isomer

- Other Topics Available(*) (546 additional topics)
 - ◆ Math and Physics (92 topics)
 - ◇ Signed fraction addition or subtraction: Basic
 - ◇ Signed fraction multiplication: Basic
 - ◇ Signed fraction division
 - ◇ Exponents and fractions
 - ◇ Introduction to inequalities
 - ◇ Introduction to exponents
 - ◇ Introduction to order of operations
 - ◇ Ordering numbers with positive exponents
 - ◇ Evaluating an expression with a negative exponent: Whole number base
 - ◇ Evaluating an expression with a negative exponent: Positive fraction base
 - ◇ Complex fraction without variables: Problem type 1
 - ◇ Introduction to square root multiplication
 - ◇ Absolute value of a number
 - ◇ Evaluating a quadratic expression: Integers
 - ◇ Combining like terms: Integer coefficients
 - ◇ Combining like terms in a quadratic expression
 - ◇ Distributive property: Integer coefficients
 - ◇ Introduction to the product rule of exponents
 - ◇ Product rule with positive exponents: Univariate
 - ◇ Introduction to the product rule with negative exponents
 - ◇ Introduction to the quotient rule of exponents
 - ◇ Simplifying a ratio of univariate monomials
 - ◇ Introduction to the power of a product rule of exponents
 - ◇ Rewriting an algebraic expression without a negative exponent
 - ◇ Squaring a binomial: Univariate
 - ◇ Multiplying binomials with leading coefficients greater than 1
 - ◇ Multiplying rational expressions involving multivariate monomials
 - ◇ Complex fraction involving univariate monomials
 - ◇ Writing a one-step expression for a real-world situation
 - ◇ Writing a multi-step equation for a real-world situation
 - ◇ Identifying solutions to a linear equation in one variable: Two-step equations
 - ◇ Identifying solutions to a linear equation in two variables
 - ◇ Additive property of equality with a negative coefficient

- ◇ Solving a multi–step equation given in fractional form
- ◇ Solving a proportion of the form $a/(x+b) = c/x$
- ◇ Solving a word problem with two unknowns using a linear equation
- ◇ Additive property of inequality with integers
- ◇ Solving a two–step linear inequality: Problem type 2
- ◇ Discriminant of a quadratic equation
- ◇ Introduction to solving a radical equation
- ◇ Solving a radical equation that simplifies to a linear equation: One radical, basic
- ◇ Graphing a line given its equation in slope–intercept form: Fractional slope
- ◇ Writing the equation of a line given the y –intercept and another point
- ◇ Finding slope given two points on a line
- ◇ Finding the slope and y –intercept of a line given its equation in the form $Ax + By = C$
- ◇ Finding x – and y –intercepts given the graph of a line on a grid
- ◇ Finding x – and y –intercepts of a line given the equation: Advanced
- ◇ Constructing a scatter plot
- ◇ Sketching the line of best fit
- ◇ Scatter plots and correlation
- ◇ Approximating the equation of a line of best fit and making predictions
- ◇ Classifying linear and nonlinear relationships from scatter plots
- ◇ Linear relationship and the correlation coefficient
- ◇ Choosing a graph to fit a narrative: Basic
- ◇ Choosing a graph to fit a narrative: Advanced
- ◇ Constructing a histogram for numerical data
- ◇ Mean of a data set
- ◇ Finding the mean of a symmetric distribution
- ◇ Population standard deviation
- ◇ Word problem involving calculations from a normal distribution
- ◇ Basic properties of logarithms
- ◇ Expanding a logarithmic expression: Problem type 1
- ◇ Expanding a logarithmic expression: Problem type 2
- ◇ Writing an expression as a single logarithm
- ◇ Evaluating an exponential function with base e that models a real–world situation
- ◇ Converting between common logarithmic and exponential equations
- ◇ Converting between natural logarithmic and exponential equations
- ◇ Solving a multi–step equation involving natural logarithms
- ◇ Solving an exponential equation by using logarithms: Decimal answers, basic
- ◇ Solving an exponential equation by using natural logarithms: Decimal answers
- ◇ Graphing an exponential function and its asymptote: $f(x) = a(e)^{x-b} + c$
- ◇ Finding an angle measure of a triangle given two angles
- ◇ Finding an angle measure for a triangle with an extended side
- ◇ Finding an angle measure for a triangle sharing a side with another triangle
- ◇ Pythagorean Theorem
- ◇ Sine, cosine, and tangent ratios: Numbers for side lengths
- ◇ Using the Pythagorean Theorem to find a sine, cosine, or tangent ratio in a right triangle
- ◇ Using a trigonometric ratio to find a side length in a right triangle
- ◇ Using a trigonometric ratio to find an angle measure in a right triangle
- ◇ Solving a right triangle
- ◇ Finding the magnitude and direction of a vector given its graph
- ◇ Finding the components of a vector given its graph
- ◇ Finding the component of a vector along another vector
- ◇ Calculating gravitational potential energy
- ◇ Using conservation of energy with gravitational potential energy
- ◇ Using conservation of energy with electrostatic potential energy

- ◇ Understanding how electrostatic force scales with charge and separation
- ◇ Understanding how electrostatic potential energy scales with charge and separation
- ◇ Calculating the magnitude of an electrostatic force using Coulomb's Law
- ◇ Understanding that electrostatic forces add as vectors
- ◇ Understanding how electrostatic energy scales with charge and separation
- ◇ Calculating electrostatic energy using Coulomb's Law
- ◆ Measurement and Matter (57 topics)
 - ◇ Multiplication of a decimal by a power of ten
 - ◇ Division of a decimal by a power of ten
 - ◇ Multiplying and dividing numbers written in scientific notation
 - ◇ Calculating positive powers of scientific notation
 - ◇ Finding negative powers of scientific notation
 - ◇ Knowing the dimension of common scientific units of measurement
 - ◇ Understanding the purpose of SI prefixes
 - ◇ Interconversion of prefixed and base SI units
 - ◇ Interconversion of prefixed SI units
 - ◇ Interconverting compound SI units
 - ◇ Interconverting derived SI units
 - ◇ Interconverting whole degree temperatures in Celsius and kelvins
 - ◇ Interconverting temperatures in Celsius and Fahrenheit
 - ◇ Addition and subtraction of measurements
 - ◇ Simplifying unit expressions
 - ◇ Multiplication and division of measurements
 - ◇ One step dosage calculations
 - ◇ Dosage calculations using patient weight
 - ◇ Dosage rate calculations
 - ◇ Rounding to a given significant digit
 - ◇ Adding or subtracting and multiplying or dividing measurements
 - ◇ Reading a measurement from an analog instrument
 - ◇ Distinguishing accuracy and precision
 - ◇ Calculating absolute and relative error
 - ◇ Naming components of the scientific method
 - ◇ Setting up a one–step unit conversion
 - ◇ Setting up a unit reprefix conversion
 - ◇ Setting up a unit conversion
 - ◇ Setting up the solution to a basic quantitative problem
 - ◇ Identifying errors in the solution to a basic quantitative problem
 - ◇ Setting up the math for a one–step quantitative problem
 - ◇ Estimating the volume in liters of a square prism object
 - ◇ Finding the side length of a cube from its volume in liters
 - ◇ Estimating the volume in liters of a spherical object
 - ◇ Estimating the volume in liters of a cylindrical object
 - ◇ Calculating volume by combining the volume of simple shapes
 - ◇ Understanding the relationship between mass, volume, and density
 - ◇ Calculating mass density
 - ◇ Using mass density to find mass or volume
 - ◇ Solving applied density problems
 - ◇ Distinguishing mixtures from pure substances through physical properties
 - ◇ Naming common laboratory separation techniques
 - ◇ Distinguishing physical and chemical properties by a macroscopic description
 - ◇ Distinguishing extensive and intensive properties
 - ◇ Distinguishing elements and compounds
 - ◇ Distinguishing compounds and mixtures

- ◇ Classifying substances from a sketch
- ◇ Distinguishing chemical and physical change
- ◇ Distinguishing solid, liquid and gas phases of a pure substance
- ◇ Using the Law of Constant Composition
- ◇ Using the Law of Multiple Proportions
- ◇ Names and symbols of important elements
- ◇ Reading a Periodic Table entry
- ◇ Understanding periods and groups of the Periodic Table
- ◇ Standard chemical and physical states of the elements
- ◇ Using the Periodic Table to identify similar elements
- ◇ Distinguishing a metal from a nonmetal by physical properties
- ◆ Atoms, Ions and Molecules (29 topics)
 - ◇ Identifying the parts of an atom
 - ◇ Counting the number of protons and electrons in a neutral atom
 - ◇ Counting protons and electrons in atoms and atomic ions
 - ◇ Finding isoprotonic atoms
 - ◇ Finding isoelectronic atoms
 - ◇ Isotopes
 - ◇ Finding atomic mass from isotope mass and natural abundance
 - ◇ Finding isotope mass or natural abundance from atomic mass
 - ◇ Counting valence electrons in an atomic ion
 - ◇ Drawing the Lewis dot diagram of a main group atom or common atomic ion
 - ◇ Counting the electron shells in a neutral atom
 - ◇ Counting the number of atoms in a formula unit
 - ◇ Writing a chemical formula given a molecular model
 - ◇ Writing a chemical formula given a chemical structure
 - ◇ Understanding the difference between a molecular and empirical formula
 - ◇ Naming binary covalent compounds
 - ◇ Predicting whether a compound is ionic or molecular
 - ◇ Distinguishing an ionic from a molecular compound by physical properties
 - ◇ Predicting the formula of binary ionic compounds
 - ◇ Naming binary ionic compounds
 - ◇ Deducing the empirical formula of a binary ionic compound from its name
 - ◇ Predicting ionic compounds formed by two elements
 - ◇ Predicting and naming ionic compounds formed by two elements
 - ◇ Identifying common polyatomic ions
 - ◇ Predicting the formula of ionic compounds with common polyatomic ions
 - ◇ Naming ionic compounds with common polyatomic ions
 - ◇ Deducing the ions in a polyatomic ionic compound from its empirical formula
 - ◇ Identifying oxoanions
 - ◇ Naming hydrates
- ◆ Chemical Reactions (44 topics)
 - ◇ Using the Avogadro Number
 - ◇ Calculating and using the molar mass of elements
 - ◇ Calculating and using the molar mass of diatomic elements
 - ◇ Calculating and using the molar mass of heterodiatomic compounds
 - ◇ Finding mole ratios from chemical formulae
 - ◇ Finding chemical formulae from a mole ratio
 - ◇ Interconverting number of atoms and mass of compound
 - ◇ Finding mass percent from chemical formulae
 - ◇ Solving applied mass percent problems
 - ◇ Elemental analysis of binary compounds
 - ◇ Elemental analysis

- ◇ Finding a molecular formula from molar mass and elemental analysis of binary compounds
- ◇ Finding a molecular formula from molar mass and elemental analysis
- ◇ Combustion analysis
- ◇ Stoichiometric coefficients
- ◇ Balancing chemical equations with noninterfering coefficients
- ◇ Writing a chemical equation from a description of the reaction
- ◇ Solving moles-to-moles limiting reactant problems
- ◇ Limiting reactants
- ◇ Understanding theoretical, actual, and percent yield
- ◇ Theoretical yield of chemical reactions
- ◇ Percent yield of chemical reactions
- ◇ Reaction sequence stoichiometry
- ◇ Calculating molarity using solute moles
- ◇ Using molarity to find solute moles and solution volume
- ◇ Calculating ion molarity using solute mass
- ◇ Dilution
- ◇ Solving for a reactant in solution
- ◇ Solving limiting reactant problems in solution
- ◇ Predicting the products of dissolution
- ◇ Predicting precipitation
- ◇ Identifying acids and bases by their chemical formula
- ◇ Determining the molar mass of an acid by titration
- ◇ Standardizing a base solution by titration
- ◇ Identifying oxidizing and reducing agents
- ◇ Identifying oxidized and reduced reactants in a metal-nonmetal reaction
- ◇ Identifying oxidized and reduced reactants in a single-displacement reaction
- ◇ Predicting whether simple electrochemical reactions happen
- ◇ Solving a redox titration problem
- ◇ Identifying combination, decomposition, single and double displacement reactions
- ◇ Identifying precipitation, combustion and acid-base reactions
- ◇ Predicting the products of a combustion reaction
- ◇ Predicting the products of a single displacement reaction involving hydrogen
- ◇ Predicting the products of a gas-evolving double displacement reaction
- ◆ Thermochemistry (20 topics)
 - ◇ Understanding how kinetic energy scales with mass and speed
 - ◇ Calculating kinetic energy
 - ◇ Using conservation of energy to predict the qualitative exchange of kinetic and potential energy
 - ◇ Calculating pressure-volume work
 - ◇ Interconverting calories and joules
 - ◇ Calculating specific heat capacity
 - ◇ Using specific heat capacity to find heat
 - ◇ Using specific heat capacity to find temperature change
 - ◇ Calculating molar heat capacity
 - ◇ Solving a basic calorimetry problem
 - ◇ Finding the equilibrium temperature when substances at different temperatures mix
 - ◇ Calculating the heat of reaction from molar reaction enthalpy and the mass of a reactant
 - ◇ Calculating heat of reaction from constant-pressure calorimetry data
 - ◇ Calculating heat of reaction from bomb calorimetry data
 - ◇ Using Hess's Law to calculate net reaction enthalpy
 - ◇ Writing a standard formation reaction
 - ◇ Calculating a molar heat of reaction from formation enthalpies
 - ◇ Solving combustion thermochemistry problems
 - ◇ Calculating the heat of reaction from bond energies and Lewis structures

- ◇ Calculating the heat of reaction from bond energies
- ◆ Electronic Structure and Chemical Bonding (49 topics)
 - ◇ Understanding the meaning of a de Broglie wavelength
 - ◇ Finding the minimum uncertainty in a position or velocity measurement
 - ◇ Interpreting the radial probability distribution of an orbital
 - ◇ Interpreting the angular probability distribution of an orbital
 - ◇ Deciding the relative energy of electron subshells
 - ◇ Deducing the allowed quantum numbers of an atomic electron
 - ◇ Calculating the capacity of electron subshells
 - ◇ Knowing the subshells of an electron shell
 - ◇ Interpreting the electron configuration of a neutral atom
 - ◇ Interpreting the electron configuration of a neutral atom in noble-gas notation
 - ◇ Writing the electron configuration of a neutral atom with s and p electrons only
 - ◇ Writing the electron configuration of a neutral atom with a filled d subshell
 - ◇ Interpreting the electron configuration of an atom or atomic ion
 - ◇ Interpreting the electron configuration of an atom or atomic ion in noble-gas notation
 - ◇ Writing the electron configuration of an atom or atomic ion with s and p electrons only
 - ◇ Identifying quantum mechanics errors in electron configurations
 - ◇ Identifying the electron added or removed to form an ion from an s or p block atom
 - ◇ Identifying the electron added or removed to form an ion
 - ◇ Understanding the definitions of ionization energy and electron affinity
 - ◇ Predicting the relative ionization energy of elements
 - ◇ Deducing valence electron configuration from trends in successive ionization energies
 - ◇ Ranking the screening efficacy of atomic orbitals
 - ◇ Understanding the organization of the electromagnetic spectrum
 - ◇ Interconverting the wavelength and frequency of electromagnetic radiation
 - ◇ Interconverting wavelength, frequency and photon energy
 - ◇ Calculating the wavelength of a spectral line from an energy diagram
 - ◇ Predicting the qualitative features of a line spectrum
 - ◇ Calculating the wavelength of a line in the spectrum of hydrogen
 - ◇ Counting bonding and nonbonding electron pairs in a Lewis structure
 - ◇ Counting electron pairs in a Lewis structure with double or triple bonds
 - ◇ Deciding whether a Lewis structure satisfies the octet rule
 - ◇ Writing Lewis structures for diatomic molecules
 - ◇ Predicting the single-bonded molecular compounds formed by two elements
 - ◇ Predicting the compound formed by two main group elements
 - ◇ Recognizing exceptions to the octet rule
 - ◇ Writing Lewis structures for an expanded valence shell central atom
 - ◇ Predicting bond polarity
 - ◇ Predicting the relative ionic character of chemical bonds
 - ◇ Predicting the arrangement of electron groups around the central atom of a molecule
 - ◇ Identifying a molecule with one central atom from its 3D shape
 - ◇ Using the AXE notation to describe a molecule with a central atom
 - ◇ Naming the shape of molecules with one central atom and no octet-rule exceptions
 - ◇ Predicting bond angles in molecules with one central atom and no octet-rule exceptions
 - ◇ Predicting and naming the shape of molecules with a central atom
 - ◇ Predicting deviations from ideal bond angles
 - ◇ Identifying hybridization in a small molecule
 - ◇ Recognizing typical LCAO molecular orbitals
 - ◇ Drawing the MO energy diagram for a Period 2 homodiatom
 - ◇ Using the MO model to predict bond order and paramagnetism
- ◆ States of Matter (48 topics)
 - ◇ Interconverting pressure and force

- ◇ Interconverting atmospheres and kilopascals
- ◇ Interconverting atmospheres and torr
- ◇ Understanding pressure equilibrium and atmospheric pressure
- ◇ Understanding Boyle's Law
- ◇ Solving applications of Boyle's Law
- ◇ Using Charles's Law
- ◇ Using the combined gas law
- ◇ Using Avogadro's Law
- ◇ Interconverting molar mass and density of ideal gases
- ◇ Calculating partial pressure of a gas from a sketch
- ◇ Calculating mole fraction in a gas mixture
- ◇ Calculating partial pressure in a gas mixture
- ◇ Calculating the mass of a gas collected over water
- ◇ Solving for a gaseous reactant
- ◇ Understanding how average molecular kinetic energy scales with temperature
- ◇ Understanding how average molecular speed scales with temperature and molar mass
- ◇ Interpreting a graph of molecular speed distribution
- ◇ Predicting how molecular speed distribution changes with temperature and molar mass
- ◇ Calculating average molecular speed
- ◇ Understanding how molecular collision rate scales with temperature and volume
- ◇ Using relative effusion rates to find an unknown molar mass
- ◇ Using thermodynamic state to order the ideality of gases
- ◇ Identifying the origin of nonideality in a gas
- ◇ Understanding the origin of the van der Waals equation of state
- ◇ Using the van der Waals equation of state
- ◇ Identifying the important intermolecular forces in pure compounds
- ◇ Calculating vapor pressure from boiling point and enthalpy of vaporization
- ◇ Calculating enthalpy of vaporization from vapor pressure
- ◇ Interpreting a Born–Haber cycle
- ◇ Calculating key distances in the fcc unit cell
- ◇ Calculating key distances in the bcc unit cell
- ◇ Finding an atomic radius from an fcc or bcc lattice constant
- ◇ Finding density from an fcc or bcc lattice constant
- ◇ Drawing a heating curve
- ◇ Using mass percent composition to find solution volume
- ◇ Calculating volume percent composition
- ◇ Finding mass or volume from percent concentration
- ◇ Calculating ionic solution composition in equivalents
- ◇ Solving applied equivalents composition problems
- ◇ Using mass concentration to find solute mass and solution volume
- ◇ Solving applied mass concentration problems
- ◇ Calculating solubility
- ◇ Using solubility to calculate solute mass or solution volume
- ◇ Understanding conceptual components of the enthalpy of solution
- ◇ Using Henry's Law to calculate the solubility of a gas
- ◇ Predicting the relative heat of hydration of ions
- ◇ Calculating and using the van't Hoff factor for electrolytes
- ◆ Kinetics and Equilibrium (10 topics)
 - ◇ Using the Arrhenius equation to calculate k at one temperature from k at another
 - ◇ Using the Arrhenius equation to calculate E_a from k versus T data
 - ◇ Calculating an equilibrium constant from a heterogeneous equilibrium composition
 - ◇ Interconverting K_p and K_c
 - ◇ Writing an equilibrium constant for a reaction sequence

- ◇ Using the small x approximation to solve equilibrium problems
- ◇ Using the van't Hoff equation to predict K at a different temperature
- ◇ Writing a complex ion formation constant expression
- ◇ Using K_f to calculate the equilibrium molarity of a complex
- ◇ Calculating the solubility of an ionic compound when a complex may form
- ◆ Acids and Bases (4 topics)
 - ◇ Predicting the acid–base properties of a binary oxide in water
 - ◇ Understanding connections between descriptions of weak acid dissociation
 - ◇ Calculating the pH of a dilute acid solution
 - ◇ Solving a polyprotic acid equilibrium composition problem
- ◆ Entropy and Free Energy (4 topics)
 - ◇ Calculating standard reaction free energy from standard free energies of formation
 - ◇ Estimating a phase transition temperature from standard thermodynamic data
 - ◇ Calculating reaction free energy under nonstandard conditions
 - ◇ Using reaction free energy to predict equilibrium composition
- ◆ Electrochemistry (5 topics)
 - ◇ Writing the half–reactions of a metal–nonmetal reaction
 - ◇ Writing the half–reactions of a single–displacement reaction
 - ◇ Designing a galvanic cell from a single–displacement redox reaction
 - ◇ Analyzing the electrolysis of molten salt
 - ◇ Calculating the mass of an electrolysis product from the applied current
- ◆ Chemistry of the Elements (38 topics)
 - ◇ Understanding main–group periodic trends in ionization energy
 - ◇ Understanding main–group periodic trends in atomic radius
 - ◇ Predicting the hydride formed by a main–group element
 - ◇ Predicting the oxide formed by a main–group element
 - ◇ Identifying a main–block group from its general properties
 - ◇ Assessing the consistency of statements relating to main–group valence electron configuration
 - ◇ Identifying Group 3A elements
 - ◇ Identifying Group 4A elements
 - ◇ Identifying Group 5A elements
 - ◇ Identifying Group 6A elements
 - ◇ Understanding the chemical formulae of interhalogens
 - ◇ Understanding how halide bond length varies down a main–block group
 - ◇ Ranking the oxidizing power of halogens
 - ◇ Interpreting an outer electron box diagram
 - ◇ Drawing the outer electron box diagram of a transition metal cation
 - ◇ Deducing the number of d electrons and unpaired spins in a transition metal cation
 - ◇ Understanding the exceptional electron configurations in the first transition series
 - ◇ Predicting the relative atomic radius of a transition metal atom
 - ◇ Predicting the relative density of a transition metal
 - ◇ Predicting the relative melting point of a transition metal
 - ◇ Predicting the highest common oxidation state of a metal in the first transition series
 - ◇ Predicting the reaction of a transition metal with a strong acid
 - ◇ Writing the formula of a metal complex from its description
 - ◇ Recognizing typical metal ligands
 - ◇ Determining the oxidation state of the metal in a complex ion
 - ◇ Naming complex cations with one type of ligand
 - ◇ Naming complex anions with one type of ligand
 - ◇ Naming complex ions
 - ◇ Determining the oxidation state of the metal in a coordination compound
 - ◇ Naming coordination compounds
 - ◇ Determining the coordination number of a metal in a complex

- ◇ Understanding the connection between geometry and coordination number of a metal complex
- ◇ Distinguishing isomers and alternate views of a metal complex
- ◇ Drawing an isomer of a metal complex
- ◇ Drawing cis and trans isomers of a metal complex
- ◇ Adding electrons to a crystal field theory energy level diagram
- ◇ Predicting color and magnetic properties from a crystal field theory energy level diagram
- ◇ Drawing a crystal field theory energy level diagram
- ◆ Organic Chemistry (76 topics)
 - ◇ Identifying hydrophobic and hydrophilic parts of an organic molecule
 - ◇ Interpreting the skeletal structure of a neutral organic molecule
 - ◇ Drawing a skeletal structure from a simple condensed structure
 - ◇ Drawing a skeletal structure from a condensed structure
 - ◇ Interpreting a skeletal structure with aromatic rings
 - ◇ Recognizing different skeletal structures
 - ◇ Understanding H atoms in a skeletal structure
 - ◇ Comparing skeletal structures related by one fewer bond
 - ◇ Using wedges and dashes in skeletal structures
 - ◇ Naming alkyl side chains
 - ◇ Understanding the basic descriptive vocabulary of hydrocarbons
 - ◇ Understanding the basic descriptive vocabulary of molecules with functional groups
 - ◇ Numbering the main chain of branched alkanes
 - ◇ Naming and drawing small alkyl substituents
 - ◇ Naming and drawing alkyl and alkoxy substituents
 - ◇ Drawing the skeletal structure of a constitutional isomer
 - ◇ Identifying a chiral molecule from its condensed structure
 - ◇ Identifying chiral centers in a cyclic molecule
 - ◇ Drawing the mirror image of a simple organic molecule
 - ◇ Identifying the enantiomer of a simple organic molecule
 - ◇ Classifying organic reactions
 - ◇ Recognizing organic acids and bases
 - ◇ Deducing oxidation state from a Lewis structure
 - ◇ Identifying oxidation and reduction in organic reactions
 - ◇ Naming branched alkanes
 - ◇ Using multiplying affixes in the names of branched alkanes
 - ◇ Naming and drawing normal alkanes.
 - ◇ Naming and drawing simple cyclic alkanes
 - ◇ Naming and drawing branched alkanes
 - ◇ Naming and drawing simple substituted cycloalkanes
 - ◇ Naming unbranched alkenes and alkynes
 - ◇ Naming alkenes and alkynes
 - ◇ Naming and drawing linear alkenes with one double bond
 - ◇ Identifying cis/trans isomerism in a small condensed structure
 - ◇ Identifying cis/trans isomerism in a skeletal structure
 - ◇ Drawing the cis or trans isomer of a small alkene
 - ◇ Naming benzene derivatives
 - ◇ Predicting the reactants or products of alkene hydrogenation
 - ◇ Predicting the reactants or products of alkene hydration
 - ◇ Naming alkyl halides
 - ◇ Naming and drawing alkyl halides
 - ◇ Naming alcohols
 - ◇ Naming and drawing alcohols without alkyl side groups
 - ◇ Naming and drawing thiols without alkyl side groups
 - ◇ Naming and drawing alcohols

- ◇ Identifying primary, secondary, and tertiary alcohols
- ◇ Identifying common alcohols from a description
- ◇ Predicting the products of symmetric alcohol dehydration
- ◇ Predicting the reactants or products of alcohol oxidation
- ◇ Understanding the common names of simple ethers
- ◇ Understanding the common names of simple ketones
- ◇ Naming and drawing ketones
- ◇ Naming and drawing aldehydes
- ◇ Predicting the reactants or products of alcohol and aldehyde oxidation
- ◇ Predicting the reactants or products of carbonyl reduction
- ◇ Identifying and drawing hemiacetals and acetals
- ◇ Predicting the reactants or products of hemiacetal and acetal formation
- ◇ Predicting the reactants or products of acetal hydrolysis
- ◇ Understanding the common names of simple amines
- ◇ Identifying primary, secondary, and tertiary amines
- ◇ Naming and drawing primary amines without alkyl side groups
- ◇ Naming and drawing secondary and tertiary amines
- ◇ Naming aldehydes and acids
- ◇ Naming and drawing carboxylic acids
- ◇ Understanding the names of carboxylate salts
- ◇ Naming and drawing unsubstituted esters
- ◇ Identifying primary, secondary, and tertiary amides
- ◇ Naming and drawing unsubstituted amides
- ◇ Understanding common names of carboxylic acids and derivatives
- ◇ Identifying positions labeled with Greek letters in acids and derivatives
- ◇ Knowing the common names of small diacids
- ◇ Predicting the reactants or products of esterification
- ◇ Predicting the reactants or products of ester hydrolysis
- ◇ Predicting the products of ester saponification
- ◇ Predicting the reactants or products of amidation
- ◇ Predicting the products of amide hydrolysis
- ◆ Biochemistry (70 topics)
 - ◇ Understanding the descriptive vocabulary of monosaccharides
 - ◇ Drawing the Fischer projection of the enantiomer of a monosaccharide
 - ◇ Drawing the Haworth projection of an aldose from its Fischer projection
 - ◇ Drawing the Haworth projection of a ketose from its Fischer projection
 - ◇ Naming and drawing cyclic monosaccharides
 - ◇ Identifying a given carbon in a cyclic monosaccharide
 - ◇ Naming and drawing the products of aldose oxidation and reduction
 - ◇ Identifying the parts of a disaccharide
 - ◇ Knowing the names and properties of common sugars
 - ◇ Identifying common polysaccharides
 - ◇ Understanding the glycosidic links in common polysaccharides
 - ◇ Understanding the basic descriptive vocabulary of fatty acids
 - ◇ Understanding lipid number notation
 - ◇ Understanding melting points trends of fatty acids
 - ◇ Identifying the components of wax esters
 - ◇ Identifying the parts of a triacylglycerol
 - ◇ Identifying the parts of a glycerophospholipid
 - ◇ Identifying the parts of a sphingomyelin
 - ◇ Identifying molecules that could be in a cell membrane
 - ◇ Recognizing the steroid nucleus
 - ◇ Classifying lipids derived from fatty acids

- ◇ Matching structure and function of common lipids
- ◇ Predicting the products or reactants of triacylglycerol hydrogenation
- ◇ Predicting the products or reactants of triacylglycerol hydrolysis or saponification
- ◇ Recognizing alpha amino acids
- ◇ Classifying amino acids
- ◇ Identifying the stereochemistry of natural amino acids
- ◇ Understanding the general acid–base properties of amino acids
- ◇ Identifying and drawing peptide bonds
- ◇ Describing peptides with 3–letter codes
- ◇ Identifying specific interactions between residues in a protein
- ◇ Predicting the location in a protein of a residue sequence
- ◇ Naming an element of protein secondary structure from a description
- ◇ Identifying changes at different levels of protein structure
- ◇ Recognizing nucleotides
- ◇ Numbering the carbons in nucleotides
- ◇ Naming and drawing nucleosides
- ◇ Naming and drawing nucleotides
- ◇ Understanding the arrangement of hydrogen bonds in DNA base pairs
- ◇ Understanding the structure of nucleic acid strands
- ◇ Writing complementary DNA sequences
- ◇ Understanding that DNA replication is semiconservative
- ◇ Identifying the major types of RNA from a description
- ◇ Understanding the relationship between DNA and mRNA base sequences
- ◇ Using the genetic code
- ◇ Classifying mutations
- ◇ Predicting reactants or products of phosphorylation
- ◇ Predicting the products of phosphoester or phosphoanhydride hydrolysis
- ◇ Understanding major biochemical energy storage and release reactions
- ◇ Identifying common redox coenzymes by their roles in a reaction
- ◇ Understanding the formation and hydrolysis of acyl–CoA
- ◇ Knowing basic facts about enzymes
- ◇ Classifying enzymes
- ◇ Identifying reactants and products from an enzyme name
- ◇ Understanding basic models of competitive and noncompetitive inhibition
- ◇ Predicting the effect of temperature or pH on enzyme activity
- ◇ Understanding the biochemistry of digestion
- ◇ Knowing inputs and outputs of the citric acid cycle
- ◇ Understanding the general mechanism of oxidative phosphorylation
- ◇ Solving citric acid cycle ATP stoichiometry problems
- ◇ Knowing inputs and outputs of glycolysis
- ◇ Completing a simplified diagram of glycolysis
- ◇ Completing a simplified diagram of glucose catabolism
- ◇ Solving carbohydrate catabolic stoichiometry problems
- ◇ Knowing the steps of beta oxidation
- ◇ Predicting the product of beta oxidation activation
- ◇ Knowing inputs and outputs of beta oxidation
- ◇ Solving fatty acid catabolic stoichiometry problems
- ◇ Predicting the products of catabolic amino acid transamination
- ◇ Completing a diagram of protein catabolism

***Other Topics Available** *By default, these topics are NOT included in this course because they are usually taught in other terms. However, they can be added using the content editor in the Instructor Module.*